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as the hydrated
salt $BaX_2 \cdot 2H_2O$,
where X is the
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hydrated
magnesium salt,
 $\text{MgSO}_4 \cdot x\text{H}_2\text{O}$,

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K_{sp}

$=[\text{Sr}^{2+}][\text{F}^-]^2$

$[\text{Sr}^{2+}] = [\text{SrF}_2]$

$2] = 8.6 \times 10^{-4}$

$[\text{F}^-] = 2[\text{SrF}_2] = 2$

$8.6 \times 10^{-4} = 17.2 \times 10^{-4}$

$K_{sp} = 8.6 \times 10^{-4}$

$17.2 \times 10^{-4} = 1.5$

$\times 10^{-10}$

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