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10.23 mol⁻¹ Specific
heat capacity of water, c
 $= 4.18 \text{ J g}^{-1} \text{ K}^{-1}$ Ionic
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 $= 1.00 \times 10^{-14} \text{ mol}^2$
 dm^{-6} at 298 K 1 tonne =
10⁶ g Arrhenius
equation: $k = Ae^{-E_a/RT}$
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